

## Nomenclature Fantabulous sheet

Write the formula for each of the following compounds.

1. dicarbon hexachloride  
 ~~$C_2F_6$~~   $C_2Cl_6$
2. silver nitrate  
 $AgNO_3$
3. Carbon monoxide  
 $CO$
4. barium nitrate  
 $Ba(NO_3)_2$
5. magnesium hydroxide  
 $Mg(OH)_2$
6. butane  
 $C_4H_{10}$
7. silver acetate  
 $AgC_2H_3O_2$
8. magnesium nitrate  
 $Mg(NO_3)_2$
9. aluminum sulfate  
 $Al_2(SO_4)_3$
10. potassium cyanide  
 $KCN$
11. iron(II) sulfate  
 $FeSO_4$
12. manganese(II) nitrate  
 $Mn(NO_3)_2$
13. chromium(III) nitrate  
 $Cr(NO_3)_3$
14. silver perchlorate  
 $AgClO_4$
15. tin(IV) chloride  
 $SnCl_4$
16. manganese(III) iodate  
 $Mn(IO_3)_3$
17. ammonium sulfide  
 $(NH_4)_2S$
18. cobalt(II) sulfate  
 $CoSO_4$
19. cobalt(II) hydroxide  
 $Co(OH)_2$
20. magnesium carbonate  
 $MgCO_3$
21. calcium oxide  
 $CaO$
22. lithium acetate  
 $LiC_2H_3O_2$
23. nickel(II) phosphate  
 $Ni_3(PO_4)_2$
24. magnesium carbonate  
 $MgCO_3$
25. cobalt(II) perchlorate  
 $Co(ClO_4)_2$
26. barium perchlorate  
 $Ba(ClO_4)_2$
27. copper(I) nitrate  
 $CuNO_3$
28. chromium(III) chloride  
 $CrCl_3$
29. tin(II) oxide  
 $SnO$
30. cobalt(II) sulfite  
 $CoSO_3$
31. Phosphoric acid  
 $H_3PO_4$
32. propane  
 $C_3H_8$
33. hydroiodic acid  
 $HI$

Name: Key

Write the name for each of the following compounds.

34.  $AgBr$  Silver Bromide
35.  $Ca_3(PO_4)_2$  Calcium phosphate
36.  $LiH$  Lithium Hydride
37.  $RaCl_2$  Radium Chloride
38.  $NaNO_3$  Sodium Nitrate
39.  $C_6H_{14}$  Hexane
40.  $Fe(NO_3)_3$  Iron(III) Nitrate
41.  $NH_4CH_3COO$  Ammonium acetate
42.  $Sr(NO_3)_2$  Strontium Nitrate
43.  $CaC_2O_4$  not needed
44.  $H_2Cr_2O_7$  dichromic acid
45.  $HCl$  Hydrochloric acid
46.  $C_2H_6$  ethane
47.  $SrI_2$  Strontium Iodide
48.  $NH_4F$  Ammonium fluoride
49.  $Ag_2S$  Silver sulfide
50.  $H_2SO_4$  Sulfuric acid
51.  $Cr_2(SO_4)_3$  Chromium sulfate
52.  $Hg(IO_3)_2$  Mercury(II) Iodate
53.  $CdBr_2$  Cadmium Bromide
54. Which of the following is not a binary compound?
  - a. potassium chloride
  - b. magnesium hydroxide
  - c. calcium bromide
  - d. carbon dioxide
55. Why is it necessary to use Roman Numerals in writing the name for the compound  $Fe(OH)_3$ ? Iron is a transition metal and has more than 1 charge.
56. Copper forms two different compounds with the chloride ion,  $CuCl$  and  $CuCl_2$ . In writing their names, how do we distinguish between them?

By using Roman numerals.

Copper(I) chloride

Copper(II) chloride