

Final Exam Review Part 2

Name: Key

Write the correct formula on the space provided for the compounds listed below

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| 1. Calcium Nitride <u>Ca_3N_2</u> | 2. Silver Oxide <u>Ag_2O</u> |
| 3. Ammonium phosphate <u>$(\text{NH}_4)_3\text{PO}_4$</u> | 4. Copper (II) Hydroxide <u>$\text{Cu}(\text{OH})_2$</u> |
| 5. heptane <u>C_7H_{16}</u> | 6. Boron trioxide <u>B_2O_3</u> |
| 7. hydro sulfuric acid <u>H_2S</u> | 8. Sodium Sulfate <u>Na_2SO_4</u> |
| 9. Cobalt (II) Carbonate <u>CoCO_3</u> | 10. Strontium hypochlorite <u>$\text{Sr}(\text{ClO})_2$</u> |
| 11. pentane <u>C_5H_{12}</u> | 12. dicarbon tetraiodide <u>C_2I_4</u> |

Write the correct name to the following compounds:

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|--|---|
| 13. $\text{Hg}(\text{IO}_3)_2$ <u>Mercury(II) Iodate</u> | 14. CaSO_3 <u>Calcium Sulfite</u> |
| 15. $\text{H}_2\text{Cr}_2\text{O}_7$ <u>Dichromic acid</u> | 16. CH_4 <u>Methane</u> |
| 17. $\text{Cu}_3(\text{PO}_4)_2$ <u>Copper(II) Phosphate</u> | 18. S_2O_3 <u>Di-sulfur tri oxide</u> |
| 19. NH_4F <u>Ammonium fluoride</u> | 20. Al_2S_3 <u>Aluminum Sulfide</u> |
| 21. FeN <u>Iron(III) Nitride</u> | 22. Na_2O <u>Sodium Oxide</u> |
| 23. HI <u>Hydro Iodic acid</u> | 24. HIO_3 <u>Iodic acid</u> |

Balance the following equations and give the type of equation in the line on the right.

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| 35. <u>2</u> $\text{C}_{12}\text{H}_{26}$ + <u>25</u> O_2 \rightarrow <u>24</u> CO_2 + <u>16</u> H_2O | 42) <u>Combustion</u> |
| 36. <u>2</u> Li + <u>2</u> H_2O \rightarrow <u>2</u> LiOH + <u> </u> H_2 | 43) <u>Single replacement</u> |
| 37. <u>2</u> Al + <u>3</u> $\text{Pb}(\text{NO}_3)_2$ \rightarrow <u>2</u> $\text{Al}(\text{NO}_3)_3$ + <u>3</u> Pb | 44) <u>Single replacement</u> |
| 38. <u> </u> Al_4C_3 + <u>12</u> H_2O \rightarrow <u>3</u> CH_4 + <u>4</u> $\text{Al}(\text{OH})_3$ | 45) <u>double replacement</u> |
| 39. <u>3</u> BaCl_2 + <u>2</u> H_3PO_4 \rightarrow <u> </u> $\text{Ba}_3(\text{PO}_4)_2$ + <u>6</u> HCl | 46) <u>double replacement</u> |
| 40. <u>2</u> I_2O_9 \rightarrow <u>2</u> I_2O_6 + <u>2</u> I_2 + <u>3</u> O_2 | 47) <u>decomposition</u> |
| 41. <u>2</u> Al + <u>3</u> Cl_2 \rightarrow <u>2</u> AlCl_3 | 48) <u>Synthesis</u> |

Prefixes: Latin: 1- mono 2- di 3-tri 4-tetra 5-penta 6-hexa 7-hepta 8- octa 9- nona 10-deca
 Organic 1-meth 2-eth 3-prop 4-but 5-pent 6-hex 7-hept 8-oct 9-non 10-dec