How to do the last question on the review:

$7020\left(\frac{1 mole}{161\_{9}}\right)\left(\frac{6}{2}\right)\left(\frac{58}{1}\right)=$ 7587 g NaCl This is the theoretical amount

Now multiply by .252, so answer is approximately 1897 grams

To answer in volume just divide the final answer of the stoichiometry problem by the density:

The final answer was 20.5 grams