

Objective B: Describe the organization of the periodic table.

Objective C: Distinguish between metals, nonmetals, and metalloids.

7. Why are groups sometimes called families.

8. Give the family name of each of the following Groups:

- a. Group 1 = Hydrogen ; Alkali Metals
- b. Group 2 = Alkaline earth metals
- c. Group 3 - 12 = Transition metals
- d. Group 16 = Oxygen family
- e. Group 17 = Halogens
- f. Group 18 = Noble gases

Study the quiz given in class. And also the power point.

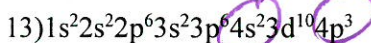
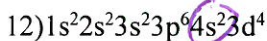
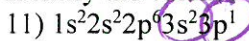
Objective D: Using a periodic table, write the electron configuration and orbital diagram for the ground state of any element in the periodic table.

9) Write the electron configuration for the elements: Na, Ar, Zn, Pt, Er

10) Draw the orbital diagrams for the following elements: Br, C, Co

Objective E: Be able to identify valence electrons in any give electron configuration.

Identify the valence electrons on the following configurations.



Objective F: Determine the number of valence electrons of a main group atom. Also be able to predict the charge of the ion that the element will form.

14) Draw the Lewis Dot structure for the following elements and predict their charge.

Na, Ca, As, Ar, P, Al, C, Cl

Vocabulary for Test: Octet rule, Orbital, Pauli exclusion principle, Aufbau principle, Hunds rule, and valence electron.

